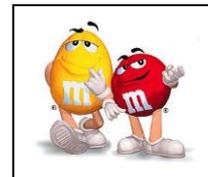


# AQA AS Chemistry-Taster lesson – Sheet2: Summer Work



Beaumont  
Chemistry

- A mole is a measure of *amount* of matter. It has a numerical value, just like the word dozen = 12; a mole =  $6.022 \times 10^{23}$  particles, quite a big number (called Avogadro's number) –here's an indication how big:
- If the whole population of the Earth ( $\approx 6.5$  billion people) were to eat 100 M&Ms *every second* it would take a mere 29 000 years!



## Using Avogadro's Number

1. Calculate the number of moles present in each of the following cases:	2. Calculate the mass of substance present in the following cases:	3. Calculate the relative molecular mass of the following substances and suggest a possible identity of each substance:
a) 2.3 g of Na	a) 0.05 moles of $\text{Cl}_2$	a) 0.015 moles, 0.42 g
b) 2.5 g of $\text{O}_2$	b) 0.125 moles of KBr	b) 0.0125 moles, 0.50 g
c) 240 kg of $\text{CO}_2$	c) 0.075 moles of $\text{Ca}(\text{OH})_2$	c) 0.55 moles, 88 g
d) 12.5 g of $\text{Al}(\text{OH})_3$	d) 250 moles of $\text{Fe}_2\text{O}_3$	d) 2.25 moles, 63 g
e) 5.2 g of $\text{PbO}_2$	e) 0.02 moles of $\text{Al}_2(\text{SO}_4)_3$	e) 0.00125 moles, 0.312 g

4. Calculate the number of particles in the following substances:

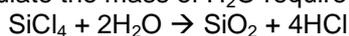
- a) 0.025 moles                      b) 2.5 g of  $\text{CO}_2$                       c) 5.0 g of Pb                      d) 100 g of  $\text{N}_2$

5. Calculate the mass of the following substances:

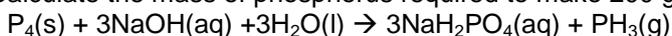
- a)  $2.5 \times 10^{23}$  molecules of  $\text{N}_2$   
b)  $1.5 \times 10^{24}$  molecules of  $\text{CO}_2$   
c)  $2 \times 10^{20}$  atoms of Mg

## Reacting Masses

6. Calculate the mass of  $\text{H}_2\text{O}$  required to react completely with 5.0 g of  $\text{SiCl}_4$ :



7. Calculate the mass of phosphorus required to make 200 g of phosphine,  $\text{PH}_3$ , by the reaction:

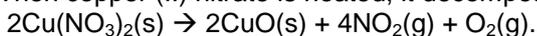


8. Lead (IV) oxide reacts with concentrated hydrochloric acid as follows:



What mass of lead chloride would be obtained from 37.2g of  $\text{PbO}_2$ , and what mass of chlorine gas would be produced?

9. When copper (II) nitrate is heated, it decomposes according to the following equation:



When 20.0g of copper (II) nitrate is heated, what mass of copper (II) oxide would be produced? What mass of  $\text{NO}_2$  would be produced?

10. A blast furnace can produce about 700 tonnes of iron a day. How much iron (III) oxide will be consumed? Assuming coke is pure carbon, how much coke would be needed to produce the necessary carbon monoxide?

